

A Three-Coordinate and Quadrupty Bonded Mo–Mo Complex

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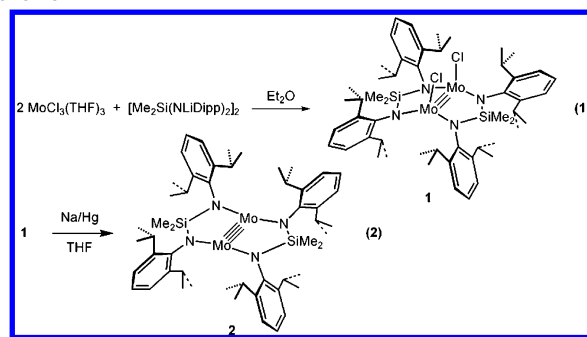
Since Professor Cotton's seminal recognition of the first dirhenium complex $[\text{Re}_2\text{Cl}_8]^{2-}$ featuring a quadruple bond between two metal atoms in 1964,¹ the field of metal–metal quadruple-bond chemistry has progressed at a dramatic rate in the past decades. Thus far, the paddlewheel motif has been one of the most prevalent structural types observed.² It is worth noting that in order to preserve the δ -interaction between two metals, the four ligands coordinating to each metal are restricted to either σ -donors or weak π -donors, such as amines, halides, phosphines, and alkoxides.^{2,3}

In addition to the regular quadrupty bonded dimetal species, coordination complexes having metal–metal multiple bonds can be stabilized by good π -donor ligands. For instance, Chisholm has pioneered the use of six amide ligands to support group VI metals in discrete “ethane-like” dimeric forms, $\text{M}_2(\text{NMe}_2)_6$ where $\text{M} = \text{Mo}$ or W .⁴ Moreover, Power's group recently reported the first isolable dichromium complex with a five-fold bond between two Cr atoms, supported by only two bulky monodentate carbyl ligands.⁵ Low-coordinate metal complexes are highly attractive due to their unsaturated coordination spheres which can serve as an efficient platform for activating small molecules such as N_2 .⁶ In this communication, we report a successful synthesis of the first symmetrically bridged⁷ and quadrupty bonded dinuclear Mo(II) amido complex of the type of M_2X_4 .

Metathesis reaction between the dimeric dilithio salts of $[\text{Me}_2\text{Si}\{\text{N}(\text{Li}(\text{Dipp})\}_2)_2$ ($\text{Dipp} = 2,6\text{-}i\text{-Pr}_2\text{C}_6\text{H}_3$)⁸ with 2 equiv of monomeric $\text{MoCl}_3(\text{THF})_3$ in diethyl ether at room-temperature resulted in the formation of the corresponding triply bonded dimolybdenum complex *syn*-1,2- $\text{Mo}_2\text{Cl}_2[\mu\text{-}\eta^2\text{-Me}_2\text{Si}(\text{NDipp})_2]_2$ **1**, (eq 1, Scheme 1) which was isolated by recrystallization from *n*-hexane in 90% yield as air- and moisture-sensitive orange crystals. The proton NMR spectrum of **1** in C_6D_6 showed four doublet resonance signals (δ 0.62, 1.29, 1.47 and 1.69) corresponding to the eight *i*-Pr groups, and two multiplet signals around δ 4.27 and 4.13 ppm which were assigned to the eight methine protons.

Crystals of **1** were analyzed by X-ray crystallography. Interestingly, in contrast to the reported staggered anti rotamers of 1,2- $\text{M}_2\text{X}_2(\text{NR}_2)_4$,^{2,9} the structure of **1** adopted the form of the eclipsed *syn* conformation, featuring two terminal chlorides and two $\text{Me}_2\text{Si}(\text{NDipp})_2$ ligands spanning a metal–metal triple bond.¹⁰ The ORTEP drawing of **1** is shown in Figure S1 (see Supporting Information). The coordination geometry of each metal center is nearly trigonal pyramidal. A striking structural feature of **1** is the large Cl–Mo–Mo–Cl torsion angle of $30.43(8)^\circ$, which is assumed as a consequence of the tetrahedral geometry around Si atoms. The Mo–N distances of 1.977(5) and 1.988(5) Å in **1** are in the range of the documented Mo–N lengths.^{2,4,9} Interestingly, despite the structural difference between **1** and complexes *anti*-1,2- $\text{MoX}_2\text{-}(\text{NMe}_2)_4$, the Mo–Mo bond length of 2.2016(10) Å for **1** is typical for Mo–Mo triple bonds, which is even slightly shorter than that

Scheme 1



in the corresponding *anti*-1,2- $\text{MoCl}_2(\text{NMe}_2)_4$,^{9b} due to greater orbital overlap between MX_3 fragments in an eclipsed ligand conformation.¹¹ Compound **1** thus provides us a good opportunity for the preparation of an unprecedented complex of the type Mo_2X_4 .

In pursuit of an unprecedented three-coordinate and quadrupty bonded Mo_2 complex, we sought to chemically reduce **1** by two electrons. The cyclic voltammogram of a solution of **1** (THF/TBAP) shows two reversible reductions at $E_{1/2} = -1.83$ and -2.01 V (relative to Fc/Fc^+) over the course of the cathodic sweep. Accordingly, reduction of an ether solution of **1** with Na/Hg gave the diamagnetic $\text{Mo}_2[\mu\text{-}\eta^2\text{-Me}_2\text{Si}(\text{NDipp})_2]_2$ **2** in 43% yield as an extremely air- and moisture-sensitive orange solid. X-ray structure analysis confirmed the dinuclear nature of **2** and the central Mo–Mo bond. The structure revealed a disorder problem, in which two molybdenum atoms were disordered over three positions. (Supporting Information) One orientation of the disordered Mo_2 in **2** depicted in Figure 1 shows a fused bicyclic skeleton containing a central Mo_2^{4+} core spanned by two ligands, and thus exhibiting a virtual C_{2h} symmetry in which the SiC_2 units lie in the horizontal mirror plane. Each Mo atom is three-coordinate by two N atoms of the two amides and one adjacent Mo atom. Atoms of N(1), Mo(1), Mo(1A), and N(1A) are coplanar, and the dihedral angle of N–Mo–Mo–N in both the five-membered SiN_2Mo_2 rings are $12.08(16)^\circ$, thus providing effective steric protection for the two Mo atoms. The bond length between Mo(1) and N(1) is 1.967(4) Å and between Mo(1) and N(2A) is 1.958(4) Å which are fairly short, and the sum of the bond angles around N(1), which adopts a trigonal-planar geometry, is 359.4° and that of N(2) is 359.8° , indicating strong π -interactions between Mo and N atoms. As is usually the case for metal–metal dimers, the most intriguing metric is the metal–metal bond length, particularly in such a low-coordinate compound as **2**. Interestingly, the bond length of Mo(1)–Mo(1A) is 2.1784(12) Å and is categorized as a long Mo–Mo quadruple bond.¹²

In addition to crystallographic data, electronic absorption and resonance Raman spectroscopy also provided insight into the existence of metal–metal quadruple bonds.² Most quadrupty bonded dinuclear complexes are vividly colored due to the small separations in energy between the δ and δ^* orbitals. Indeed, an electronic

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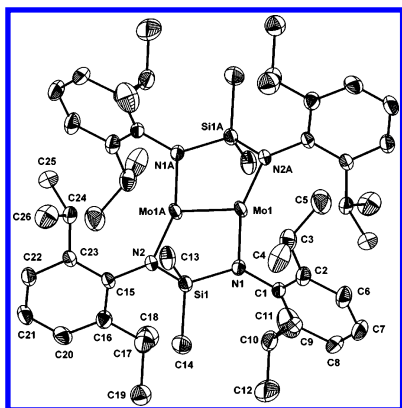


Figure 1. Molecular structure of **2** (thermal ellipsoids at the 30% probability level). Selected bond distances (Å) and angles (deg): Mo(1)–Mo(1A), 2.1784(12); Mo(1)–N(1), 1.967(4); Mo(1)–N(2A), 1.958(4); Si(1)–N(1), 1.747(4); Si(1)–N(2), 1.765(4); N(1)–Mo(1)–Mo(1A), 97.16(11); N(2)–Mo(1A)–Mo(1), 99.37(11); N(1)–Mo(1)–N(2A), 159.56(16); N(1)–Si(1)–N(2), 104.31(19); Si(1)–N(2)–Mo(1), 114.5(2); Si(1)–N(1)–Mo(1), 113.2(2).

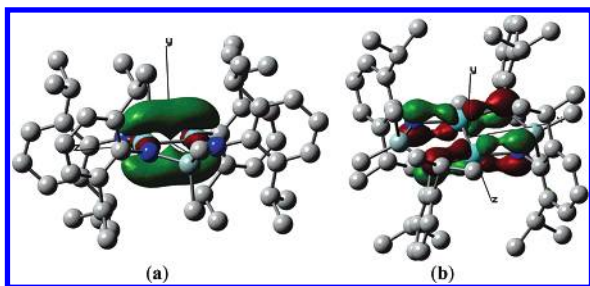


Figure 2. Contour plots of HOMO (a) and HOMO–11 (b) of **2**.

absorption band of the orange-colored complex **2** was observed at about 17240 cm^{-1} , which is assigned tentatively to an allowed transition from the molecule's δ^2 ($1A_g$) to its $\delta\delta^*$ ($1B_u$) electronic state, assuming idealized C_{2h} symmetry for the complex. A time-dependent density functional theory (TDDFT) calculation of the electronic transition gave a value well in accord with the experimentally observed quantity (Table S2, Supporting Information). Furthermore, the totally symmetric metal–metal stretch, $\nu(\text{Mo}–\text{Mo})$, was found at the frequency of 343 cm^{-1} in the resonance Raman spectrum falling in the range of frequencies between 330 and 430 cm^{-1} for the documented quadruply bonded dimolybdenum species.²

With the intent of gaining an understanding of the electronic structure and bonding of **2**, we carried out electronic structure computations using density functional theory at the BLYP level. It is noteworthy that sterically encumbered ligand variants are necessary to stabilize **2**. Attempts to model **2** via replacing Dipp groups with H atoms failed due to the crash of the molecular shape. Consequently, calculations on authentic formula were engaged, and the optimized geometry was constrained to C_{2h} . The computed Mo–Mo bond length is 2.168 Å, which agrees very well with the experimental Mo–Mo value of 2.1784(12) Å. As for the electronic structure, attention should be paid to $d_{x^2-y^2}$ and d_{xy} orbitals¹³ with two Mo atoms being defined to lie on the z axis. The $d_{x^2-y^2}$ orbitals of Mo atoms in **2** are used for δ -bond formation (HOMO, Figure 2a), although half of each $d_{x^2-y^2}$ orbital is engaged in the formation of two Mo-to-N σ -bonds, while two d_{xy} orbitals are used to form four Mo-to-N π -bonds (HOMO–11, Figure 2b). This is in contrast to the bondings between two metals in conventional paddlewheel structural motifs in which overlap of two d_{xy} orbitals gives rise to a δ -bond in paddlewheel structures while each metal uses $d_{x^2-y^2}$ to form metal-to-ligand σ -bonds. Moreover, not only does the contour

plot of HOMO have contribution from $d_{x^2-y^2}$ (69.61%), it also contains 30.30% of s orbital and 0.09% of p orbital on the basis of natural bond orbital (NBO) analysis.¹³

The chemical property of **2** is consistent with the observation made from electrochemical measurements (i.e. the oxidative addition).¹⁴ For example, exposure of **2** to organic chlorides, such as CH_2Cl_2 or 1,2- $\text{C}_2\text{H}_4\text{Cl}_2$, quickly gives **1**, which can then be converted once again to **2** upon reduction. The result provides us the opportunity to further explore the potentially rich chemistry of inter-conversion between M_2X_6 triple bonds and M_2X_4 quadruple bonds.

In summary, we have prepared an unusual triply bonded dimolybdenum complex, *syn*-1,2- $\text{Mo}_2\text{Cl}_2[\mu-\eta^2-\text{Me}_2\text{Si}(\text{NDipp})_2]_2$, **1**, from which the first three-coordinate and quadruply bonded dimolybdenum complex $\text{Mo}_2[\mu-\eta^2-\text{Me}_2\text{Si}(\text{NDipp})_2]_2$ **2** can be isolated upon reduction of **1**. Complex **2** exhibits an electronic structure different from that of the conventional paddlewheel structures. Reactivity studies of **2** are underway.

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Supporting Information Available: Experimental details for the synthesis of **1** and **2**, cyclic voltammetry, UV–vis, X-ray crystallographic data, including tables and CIF files, and details of the computational study (DFT). This material is available free of charge via the Internet at <http://pubs.acs.org>.

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